Academic Year العام الدراسي	2023/2024
Term الفصل	3
Subject المادة	Chemistry /INSPIRE الكيمياء / المسير
Grade الصف	11
PLAN -M	
Stream المسار	Advanced المتقدم
Number of MCQ عدد الأسئلة الموضوعية	20
Marks of MCQ درجة الأسئلة الموضوعية	5
Number of FRQ عدد الأسئلة المقالية	0
Marks per FRQ الدرجات للأسئلة المقالية	0
Type of All Questions نوع كافة الأسئلة	الأسئلة الموضوعية /MCQ
Maximum Overall Grade الدرجة القصوى الممكنة	100
Exam Duration - مدة الامتحان	150 minutes
Mode of Implementation - طريقة التطبيق	SwiftAssess
Calculator الآلة الحاسبة	Allowed مسموحة

Question* السؤال*		Learning Outcome/Performance Criteria** نتائج التعلم/ معايير الأداء**	Reference(s) in the Student Book (English Version) المرجع في كتاب الطالب (النسخة الإنجليزية)	
			Example/Exercise مثال/تمرين	Page الصفحة
MCQ - الأسئلة الموضوعية	1	CHM.5.3.05.001.01 Distinguish between oxidation and reduction in terms of loss and gain of electrons	**Text book +fiqure 1**	**156 and 157**
	2	CHM.5.3.05.001.10 Write oxidation-half reaction and reduction-half reaction for a redox reaction	**Text book +fiqure 2** Example Problem1 + Practice problems no.1	**157and 161**
	3	CHM.5.3.05.001.08 Identify oxidizing agent and reducing agent in a redox reaction	**Text book** Example Problem1 + Practice problems no.2&3	**159 and 161**
	4	CHM.5.3.05.001.03 Assign oxidation number to atoms, ions and compounds according to a set of rules	**Example Problem2 + Practice problems**	**162,163**
	5	CHM.5.3.05.001.04 Distinguish between oxidation and reduction in terms of change in oxidation number	**Text book**	**164**
	6	CHM.5.3.05.001.07 Define oxidizing agent and reducing agent in a redox reaction	**Text book +table 2**	**159**
	7	CHM.5.3.05.002.01 Describe the steps for balancing redox reactions, in acidic medium, by the half-reaction method	**Text book +table 5**	**169**
	8	CHM.5.3.05.002.03 Balance redox reaction in acidic medium using half-reaction method	**Example Problem5+ Practice problems**	**171**
	9	CHM.5.3.05.002.03 Balance redox reaction in acidic medium using half-reaction method	**Text book +table 6**	**169 and 170**
	10	CHM.5.3.05.002.05 Balance redox reaction in basic medium using half-reaction method	**Example Problem5+ Practice problems no.25**	**171**
	11	CHM.5.3.05.007.01 Describe an electrochemical cell while specifying its types	**Text book + fiqures 1**	**178**
	12	CHM.5.3.05.007.02 Identify components of a voltaic or galvanic cell (anode, cathode, salt bridge or porous barrier, wires, electrolyte compartments); while explaining the role of each component, when does the reaction start and determining the direction of electron and current flow	**Text book + fiqure 2**	**179**
	13	CHM.5.3.05.004.02 Describe standard hydrogen electrode (SHE), while identifying the importance of its E° value and writing the half-cell reactions of the two possible reactions that could occur at the hydrogen electrode	**Text book + fiqure 5**	**182**
	14	CHM.5.3.05.004.03 Define the reduction potential and standard electrode potential (E°)	**Text book +table 1**	**183**
	15	CHM.5.3.05.004.02 Describe standard hydrogen electrode (SHE), while identifying the importance of its E° value and writing the half-cell reactions of the two possible reactions that could occur at the hydrogen electrode	**Text book + fiqure 6**	**184**
	16	CHM.5.3.05.007.05 Use the half-cell standard reduction potentials to calculate the electrochemical cell standard potential	**Text book + Example Problem1 + Practice problems no.1-4**	**185 , 186 and 187**
	17	CHM.5.3.05.007.04 Write the cell notation and the overall chemical equation for a redox reaction occurring in a voltaic cell	**Text book**	**184,185**
	18	CHM.5.3.05.008.01 Use the standard reduction potentials to predict if a reactionspontanous or nonspontanous	**Text book + Practice problems no.5-9**	**187 and 188**
	19	CHM.5.3.05.011.03 Compare between electrolytic cell and voltaic cell in terms of identifying where will reduction and oxidation processes take place, anode, cathode, direction of electron flow and current flow and spontaneity of the reaction occurring	**Text book + fiqure 18**	**200**
	20	CHM.5.3.05.011.05 Define electroplating while describing how it works, identifying anode, cathode and electrolyte needed for an electrolytic cell in which a selected metal is to be plated on an object (car or spoon,…etc)	**Text book + fiqure22**	**204,205**

*	**Questions might appear in a different order in the actual exam**
*	قد تظهر الأسئلة بترتيب مختلف في الامتحان الفعلي
**	**As it appears in the textbook(UAE Edition Grade 12 Aavance Student Edition) , LMS, and (Main_IP).**
**	كما وردت في كتاب الطالب(كتاب الطلاب الصف الثاني عشر المتقدم طبعة دولة الامارات العربية المتحدة) و LMS والخطة الفصلية .