

Lesson Check: Oxidation and Reduction

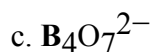
1) The most electronegative atom in a compound has a charge that is _____.

- ☐ zero
- ☐ positive
- ☐ neutral
- ☐ negative

2) Which of the elements is the strongest oxidizing agent?

- ☐ oxygen
- ☐ iodine
- ☐ fluorine
- ☐ chlorine

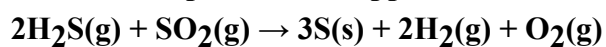
3) Determine the oxidation number of the boldface element in these ions.



4) Explain how the sulfite ion (SO_3^{2-}) differs from sulfur trioxide (SO_3).

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- 5) In a reaction, more than one element can be oxidized or reduced. In terms of oxidation and reduction, explain what happens in the following reaction.**



1) Determine the number of atoms in 3.54 mol Mn.

[illegible]

Which of the following counting units would be appropriate for counting:

g	g	P
_____	_____	_____
sheets of paper sold at a store?	_____	(Blank 1)
_____	_____	_____
atoms and molecules?	_____	(Blank 2)
_____	_____	_____
cookies?	_____	(Blank 3)
_____	_____	_____
gloves?	_____	(Blank 4)

- dozen
- ream
- mole
- pair

- dozen
- ream
- mole
- pair

- dozen
- ream
- mole
- pair

- dozen
- ream
- mole
- pair

3) Explain how measuring 6.00 g of carbon-12 indirectly counts one-half Avogadro's number of carbon-12 atoms.

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Lesson Check: Measuring Matter

4) Determine the number of representative particles in each substance.a. 4.45 mol of $\text{C}_6\text{H}_{12}\text{O}_6$ b. 0.250 mol of KNO_3 c. 2.24 mol of H_2

d. 9.56 mol of Zn

5) Determine the number of moles in each substance.a. 3.25×10^{20} atoms of leadb. 4.96×10^{24} molecules of glucosec. 1.56×10^{23} formula units of sodium hydroxided. 1.25×10^{25} copper(II) ions